Analysis of CO2 frost formation properties in cryogenic capture process

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Analysis of CO$_2$ frost formation properties in cryogenic capture process

Chun-Feng Song $^1$, Yutaka Kitamura $^{1*}$, Shu-Hong Li $^1$, Wei-Zhong Jiang $^2$

$^1$ Graduate School of Life and Environmental Sciences, University of Tsukuba, 1-1-1, Tennodai, Tsukuba, Ibaraki 305-8572, Japan

$^2$ Key Laboratory of Agricultural Engineering in Structure and Environment, Ministry of Agriculture, China Agricultural University, Qinghua Donglu 17, Beijing 100083, China

* Corresponding author. Tel.: +81 0298-53-4655; Fax: +81 0298-53-4655.

E-mail address: kitamura.yutaka.fm@u.tsukuba.ac.jp
Abstract

With the development of CO₂ capture technologies, the cryogenic approach provides a promising alternative for greenhouse effect mitigation. In previous work, a novel CO₂ capture process was developed based on Stirling coolers (SC). However, the properties of frosted CO₂ that deposits onto the heat exchanger of SC are not completely understood but it plays an important role as a heat transfer medium and affects subsequent CO₂ deposition. In order to improve the system performance and capture efficiency, a thorough analysis of the CO₂ frost process becomes especially significant. Based on this aim, numerical and experimental analysis of CO₂ frost formation in the cryogenic capture process has been investigated in the present work. The results show that the frosted CO₂ layer significantly influences the heat transfer between the gas stream and cooling fins. Over time (from 0 to 60 min), the thickness of the frost layer increased from 0 to 3.0 mm. The thermal conductivity increased from 0 to 0.4 W/(m·K). When the flow rate was set at 1 L/min, the temperature of the cold head of the SC varied from -105.2 °C to -102.1 °C. When the flow rate was 3 L/min, the temperature rose from -106.3 °C to -98.0 °C.

*Keywords:* CO₂ capture, cryogenic, frost layer, thickness, thermal conductivity
### Nomenclature

<table>
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<th>Symbol</th>
<th>Description</th>
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<tr>
<td>$T_a$</td>
<td>Temperature of ambient, °C</td>
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<tr>
<td>$T_g$</td>
<td>Temperature of flue gas, °C</td>
</tr>
<tr>
<td>$T_f$</td>
<td>Temperature of frost layer, °C</td>
</tr>
<tr>
<td>$t_1$</td>
<td>Idle operating time, min</td>
</tr>
<tr>
<td>$t_2$</td>
<td>Capture time, min</td>
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### Greek letters

<table>
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<tr>
<th>Symbol</th>
<th>Description</th>
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<tr>
<td>$v$</td>
<td>Flow rate of flue gas, L/min</td>
</tr>
<tr>
<td>$\lambda_c$</td>
<td>Thermal conductivity for individual unit volume, W/(m·K)</td>
</tr>
<tr>
<td>$\lambda_f$</td>
<td>Effective thermal conductivity of frost layer, W/(m·K)</td>
</tr>
<tr>
<td>$\rho_f$</td>
<td>Density of frost, kg/m³</td>
</tr>
<tr>
<td>$\omega_a$</td>
<td>Humidity of ambient, %</td>
</tr>
<tr>
<td>$\delta_f$</td>
<td>Thickness frost layer, mm</td>
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### Abbreviations

<table>
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<tr>
<td>CCS</td>
<td>CO₂ capture and storage</td>
</tr>
<tr>
<td>SC</td>
<td>Stirling cooler</td>
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<tr>
<td>FGD</td>
<td>Flue gas desulphurization</td>
</tr>
<tr>
<td>ESP</td>
<td>Electrostatic precipitators</td>
</tr>
<tr>
<td>SCR</td>
<td>Selective catalytic reduction</td>
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1. Introduction

CO2 capture and storage (CCS) is proposed as an important strategy to help mitigate the increasing serious issue of global warming. CCS consists of three steps: capture, transport and storage. Of these three steps, CO2 capture has attracted the most attention due to the technical challenges associated with its implementation. Currently, the main CO2 capture technologies contain chemical absorption, membrane separation, cryogenic distillation and physical adsorption. Among these technologies, amine based absorption is the most technically mature and commercially viable to be integrated into coal-fired power plants (Liang et al., 2011). The main advantages of this technology are that it can be retrofitted to existing power plants and other industrial stations with minimum modification. Furthermore, amine scrubbing is suitable for low CO2 partial pressure in flue gas (Luis et al., 2012). However, there are still some drawbacks: 1) the energy consumption required for regeneration; 2) the corrosion of the installations; 3) the environmental impact (Davison, 2007; Huang et al., 2011). For this reason, various research has been carried out on the improvement of existing technologies and the exploitation of novel processes.

Cryogenic CO2 separation technology is attracting increased attention due to its significance in recent years. As a promising alternative, this technology has several advantages: 1) to avoid the energy penalty in the regeneration process; 2) to prevent
the issue of corrosion caused by solvent; 3) to realize a CO$_2$ capture method that is environmental friendly without causing degradation. In previous work (Song et al., 2012a), a novel CO$_2$ capture process based on Stirling coolers (SC) was exploited. The whole system is chilled by SCs, and the moisture and CO$_2$ in flue gas were captured at different places. Simultaneously, the residual gas (such as N$_2$) passes through the system without phase change. The core part of the process is that CO$_2$ in the flue gas is captured in solid form. When exposed to the heat exchanger under the frost point of CO$_2$, the CO$_2$ in the gas stream will anti-sublimate and frost on the surface of the cold head. With the purpose of enhancing capture efficiency, intensive study on CO$_2$ frost properties is an imperative task.

Frost formation is a common phenomenon that exists in various fields, such as cryosurgery, meteorology, agriculture, aerospace industry, refrigeration industry, civil engineering (Lee and Ro, 2005). The investigation of the frost process in different engineering processes has been widely implemented (Seker et al., 2004; Kim et al., 2009; Schmelzer, 2003). It is worth noting that in spite of this a relatively large amount of literature has investigated the frost formation process, however most relate to moisture and there are still gaps and inconsistencies (O’Neal and Tree, 1985). Reports with respect to the frosting process of CO$_2$ in the greenhouse gas control field are few (Shchelkunov et al., 1986; Sumarokov et al., 2003). According to Ogunbameru et al. in
1973, the frost behavior of CO$_2$ is different from moisture. In the moisture frost process, the density increases continuously with the total mass of the deposition (Na and Webb, 2004a). By contrast, the initial increase rate in density for the frosted CO$_2$ is large but in the later stages it decreased (Brian et al., 1969). Meanwhile, the thermal conductivity of solid CO$_2$ is related to the variation of the layer density (Cook and Davey, 1976). Therefore, the frost formation of CO$_2$ is an intricate deposition process. In light of these characteristics, it is necessary to investigate the CO$_2$ frost formation process in order to improve cryogenic CO$_2$ capture performance.

The objective of this work is to numerically analysis the properties of the CO$_2$ frost formation process by the SC based cryogenic CO$_2$ capture technology. In order to achieve this aim, the influence of the frost layer on thermal conductivity and temperature variation has also been investigated in detail. Moreover, the variation of the frost layer thickness has been simulated.

The paper is organized as follows: First, the base case of CO$_2$ capture is described. Then the whole frost formation process is explained in three sections. Subsequently, the experimental conditions and the structure of the heat exchanger for CO$_2$ frost are introduced. Finally, the variation of frost layer (thickness and thermal conductivity) and temperature in the capture process is simulated and analyzed in detail.

2. Base case
The schematic of the exploited CO₂ capture process is shown in Fig. 1. The whole capture process can be described as follows: Prior to CO₂ recovery, the gas stream from the combustion furnace flows through the removal units. There the acid gases (such as SOₓ and NOₓ) and particulate can be separated as they affect the capture performance and installation. SOₓ removal is usually achieved in a flue gas desulphurization (FGD) unit. Particulate matter such as fly ash is removed by electrostatic precipitators (ESP). NOₓ is removed using selective catalytic reduction (SCR). Then, the gas mixture is fed to the developed cryogenic capture system which is mainly composed of three SCs. By the first SC, the flue gas is chilled to (between 40 and 50 °C) from a high initial temperature (around 60 °C after desulphurization). The function of this section is to decrease the mechanical shock of the gas stream to the installation. It should be pointed out that as the temperature of the flue gas decreases rapidly to the freezing point of H₂O, the moisture in the gas condenses and is recovered at the condensate water outlet. After the pre-chilling step, the gas stream is introduced into the CO₂ separation tower. In this section, the CO₂ in the flue gas frosts onto the heat exchanger of the SC-2 under the cryogenic condition (around -105 °C) and the residual gas (mainly as N₂) is exhausted without phase change. In this way, the CO₂ is recovered from flue gas. In the last step, the captured CO₂ is stored in the column chilled by SC-3 to avoid sublimation. Furthermore, since the temperature of
residual gas and condensate water is low, it can be used to cool down the follow-up
gas stream by heat exchanger for energy conservation. In addition, the flue gas in our
research is simulated by synthetic gas (N₂/82 vol.% , CO₂/13 vol.% and H₂O/5 vol.%).
After the cryogenic CO₂ capture process, the composition of the residual gas is
N₂/99.38 vol.% and CO₂/0.62 vol.% . Furthermore, the average flow rates of the initial
and final gas stream are 3 L/min and 2.47 L/min, respectively. The details of the
capture process can be found in previous work (Song et al., 2012a).

3. CO₂ frost formation process

As mentioned in the relevant research, the whole CO₂ frost formation process can
be divided into three periods (Hayashi et al., 1977). 1) Crystal growth period. 2) Frost
growth period. 3) Frost formation period. The sketch of the frost formation process is
depicted in Fig. 2.

3.1 Crystal growth period

This is a relatively short period during the early stage of the frosted CO₂ formation
process. At the beginning, the CO₂ in the flue gas anti-sublimates into small crystals
and is simultaneously extracted from the gas stream by frosting onto the cooling fin of
the heat exchanger. The frosted CO₂ plays a role as the nucleus for further
crystallization. New CO₂ crystal grows on the previous nucleus and spreads in all
directions. It needs to be pointed out that this period has a significant influence on the further frost growth process. However, the characteristics of this period are complexity and instantaneity. Therefore, the effect of this period on the whole capture performance is taken into account by setting appropriate initial parameters (Cui et al., 2011).

3.2 Frost growth period

The frost growth period is characterized by a porous frost structure growing. The structure consists of solid CO\(_2\) (99.89 mol.%), a small quantity of residual gas (around 0.04 mol.%) and a trace of H\(_2\)O (approximate 0.07 mol.%). Thus, in this period, the frost layer becomes visibly rugged due to new crystal formation on the surface. The subsequent CO\(_2\) crystal leads to further growth and densification of the frost layer. The thickness of the frost deposition increases rapidly.

3.3 Frost densification period

In the last period, the frost formation is fully developed. The increasing rate of deposition slows and the density increases gradually. As the frost layer grows, the heat resistance of the frost layer also rises. This reflects in the decrease in the effective thermal conductivity of the cold head and causes a surface temperature increase. When the temperature of the frost layer surface becomes higher than the freezing point, the
CO₂ in the subsequent gas stream cannot frost. Meanwhile, the following feed gas can cause that part of the frosted CO₂ on the layer surface to sublimate. The sublimation results in a thinning layer, and then enhances the thermal conductivity of the layer. The CO₂ in the gas stream frosts again until a dynamic equilibrium of anti-sublimation and sublimation of the whole frost process is achieved.

4. Materials and methods

4.1 Materials

In order to measure the temperature distribution, eight K-type thermocouples were placed on different places in the system, and all the thermocouples are connected to a data-acquisition system. Commercial software is used to record and process data in a database format using a computer. The temperature data measured by the thermocouples is recorded every 0.1 s during the total testing time. The ambient temperature ($T_a$) and humidity ($\omega_a$) are measured with the thermo hygrometer in the laboratory. All the heat exchangers in the system are made from copper. The structure of the heat exchanger of SC-2 is shown in Fig. 4(a). The details of its size are described in Fig. 4(b). The length and diameter of the cold head are 230.0 and 36.0 mm, respectively. The length and width of the cooling fin are 15.0 and 1.0 mm, respectively. The interval between the fins is 6.0 mm.
4.2 Methods

The formation of the frost layer has a significant influence on the performance of the whole capture process. In order to investigate the variation of the frost layer in the CO$_2$ capture process, two parameters are considered to represent the characteristics of variation, namely thickness ($\delta_f$) and thermal conductivity ($\lambda_f$).

The thickness of the frost layer is a significant parameter that reflects the properties of the frost process. Its variation process is observed by use of a charge coupled device (CCD) camera. Simultaneously, tick marks are added among the cooling fins of the cold head to facilitate thickness measurement. It is obvious that the deposited frost layer will lead to an increase in thermal resistance. Meanwhile, it influences the heat flux between the cold head of SC-2 and the gas stream, and results in the decrease of capture efficiency. Thus, in order to evaluate the characteristic of the frost layer, the investigation of effective thermal conductivity ($\lambda_f$) is necessary. The detail calculation method is described as follows (Lenic et al., 2009):

$$\lambda_f = \frac{\delta_f}{\sum_{c=1}^{\delta} \frac{1}{\lambda_c}}$$

(1)

where $\delta_f$ represents the thickness of the frost layer. $\lambda_c$ represents the coefficient of the thermal conductivity for individual unit volume.
4.3 Model of CO$_2$ frost process

According to previous discussion, knowledge of the characteristics of CO$_2$ frost is important to improve the capture performance and decrease the energy penalty of the cryogenic system. In order to simplify the intricate CO$_2$ frost process, some assumptions are mentioned as follows: 1) the water in the frost layer is neglected due to its fractional content. 2) the density of the frost layer on the heat exchanger is homogeneous. 3) radiation effect within the frost layer is negligible.

The CO$_2$ frost process on the cold surface is simulated in Fig. 3. The mass and heat transfer processes have been introduced in our previous work (Song et al., 2012b). The thermal dynamic behavior of the frosted CO$_2$ layer is described using empirical models. The summaries of the frost density ($\rho_f$) and thickness ($\delta_f$) correlation model are listed in Tables 1 and 2. The correlations published in the literature for thermal conductivity ($\lambda_f$) calculations are shown in Table 3.

5. Results and discussion

5.1 Frost layer

5.1.1 Growth of the frost layer

When the temperature dropped to the freezing point of CO$_2$, the flue gas was loaded
into the system. Initially, there was no frost present on the heat exchanger (as shown in Fig. 5a). As time went on, the frost began to form, temperature increased and the crystals of frosted CO$_2$ became gradually larger (as shown in Fig. 5b). Then, as the crystal of solid CO$_2$ gathered, frost grew on the heat exchanger and its volume expanded rapidly (as shown in Fig. 5c). At the full growth stage, the captured CO$_2$ almost totally covered the cooling fin of the heat exchanger (as shown in Fig. 5d).

The effect of the gas flow rate on the thickness of the frost layer is shown in Fig. 6. The result indicates that the thickness of the frost layer increased continuously with time however the rate of deposition was not constant. Initially, the growth rate was slow and then increased rapidly. That can be explained by the fact that during the stage of crystal growth (described in section 3.1), the variation in deposition was moderate. When the CO$_2$ crystal covered the cooling fin, the growth rate of the frost layer accelerated. In the frost formation period, the frost layer increased sufficiently. Meanwhile, the rate of frost deposition decreased again and eventually approached zero. In this period, the variation of the frost layer mainly presented in density. After 60 minutes, the thickness of frost layer reached 3.0 mm. Additionally, it can also be seen that along with the increase in the flow rate, the thickness of the frosted CO$_2$ grew accordingly.

5.1.2 Thermal conductivity of the frost layer
The influence of the flow rate for the gas stream on the thermal conductivity ($\lambda_f$) of the frost layer is shown in Fig. 7. From the results, it indicates that the thermal conductivity increased with capture time passed. At the beginning of the capture process, the effective conductivity of the frost layer increased quickly, and with time decelerated. In addition, with the increase in the flow rate of feed gas, the increase rate of the thermal conductivity obviously accelerated. It is worth noting that after 20 minutes the variation of the thermal conductivity of the frost deposition amplified slowly, and reached around 0.4 W/(m·K) after 60 min. This can be explained by the fact that the rise in thermal conductivity with time is due to the variation of the thickness and density of the frost layer. At the initial period (from 0 to 20 min), the CO$_2$ crystals deposited on the surface of the cold head and the growth rate of frost layer was fast. Consequently, the thermal conductivity increased and the rate was also fast. At a subsequent time (from 20 to 60 min), more CO$_2$ crystal groups formed around the existing crystals and the gas pores were filled up. The solid phase became denser through internal diffusion. However, the variation of the thickness for the frost layer slowed down. Therefore, the increase rate in the thermal conductivity decelerated.

5.2 Temperature variation

As the frosted CO$_2$ layer grows, a gradual increase in the frost surface temperature
occurs due to the augmentation of heat resistance which is unfavorable to the capture process. This is because the temperature driving force is reduced due to the frost insulation on the heat exchanger surface, and hence energy and mass transfer rates drop. As a consequence, this results in a decrease of CO₂ capture efficiency.

Fig. 8 indicates that SC-1 has a significant influence on the temperature of SC-2, and impacts the whole capture performance. During the capture process, SC-1 played a part in pre-chilling the flue gas and preventing the rapid increase of the temperature in the capture tower. The temperature of SC-2 is lower when SC-1 is working compared to stop.

Fig. 9 indicates the influence of flue gas flow rate on the temperature of SC-2. At the initial stage of gas inflow, the temperature of the heat exchanger of SC-2 increased quickly. As the frost layer formed on the surface of the cold head, the increase in speed trended to constant. From the results, it can be concluded that if flow rate was controlled slightly, the lower the temperature in the system can be maintained. This is because a high flow rate would lead to a high increase speed in the CO₂ frost layer. Due to the frost point of CO₂ being far above the temperature of the heat exchanger, when the frost formed, the temperature of the heat exchanger increased rapidly. Along with the increase in flow rate, the temperature change also increased gradually.

The influence of temperature variation of SC-1 on the properties of the frost layer is
investigated in Fig. 10 and 11. In Fig. 10, with the decrease in the temperature of SC-1, the rate of frost growth increases. When the temperature of SC-1 was set to -20 °C, the frost deposited slowly during the initial 20 minutes. Then the layer grew rapidly in the following 20 minutes. Finally, the rate of deposition decreased again and tended to zero. This is because during the initial 20 minutes, the CO₂ formed crystals on the surface of the cold head. Therefore, the frost layer changed slightly. From the 20 minute mark, the CO₂ crystal fully covered the surface of the cold head and the density of frost layer grew rapidly. When the frost layer had increased sufficiently, the rate decreased again from the 40 minutes mark. With the decreasing temperature of SC-1, the crystal formation stage became shorter and the rate of deposition faster. When the temperature dropped down to -60 °C, the variation of frost thickness stabilized after 30 minutes. In Fig. 11, at the initial period the thermal conductivity was almost zero due to little frosted CO₂ on the cold head. While capture time passed, the heat transfer coefficient of the frost layer increased rapidly but the increase rate decelerated due to the decrease in the densification rate. After 60 minutes, the thermal conductivity of deposited CO₂ was around 0.4 W/(m·K). Meanwhile, with the decrease in temperature of SC-1, the thermal conductivity of the deposition increased. That is because the lower temperature of SC-1 facilitates the pre-chilling of the flue gas, and CO₂ solidification onto the cold surface.
6. Conclusion

Frost formation is a vital consideration in cryogenic CO$_2$ capture technology, and the phenomenon of CO$_2$ frosts on the surface of the heat exchanger has an important influence on capture performance. Based on this, the growth processes of the CO$_2$ frost layer and its effect on temperature variation have been numerically analyzed in this paper. The effect of frost layer growth on thermal conductivity and temperature variation was also investigated. According to the results of simulation and experiments, the following conclusions are given:

- The thickness of the frost layer increased with time but the rate of CO$_2$ deposition decreased gradually. Furthermore, with the increase in the flow rate, the frost growth also accelerated. After 60 minutes, the thickness of frost layer was 3.0 mm.

- The thermal conductivity of the frost layer increased due to the growth of the frost layer. Nevertheless, as time passed, its increase rate reduced. After 60 minutes, the thermal conductivity of the frost layer increased to around 0.4 W/(m-K).

- The frost surface temperature increased rapidly as frost insulated the heat exchanger. Meanwhile, with the decrease in the flow rate, the increase rate of temperature decelerated. When the flow rate was set at 1 L/min, the temperature
of the cold head for the SC varied from -105.2 °C to -102.1 °C after 60 minutes. When the flow rate was 3 L/min, the temperature rose from -106.3 °C to -98.0 °C. However, the rate of the increase declined and approached zero when the energy balance between the gas stream and frost layer was reached.

Acknowledgement

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carbon dioxide from binary gas mixtures. Translated from Inzhenerno-Fizicheskii Zhurnal 51 (6), 965-970.


**Figures Captions:**

Fig. 1 Schematic diagram of the SC based cryogenic CO$_2$ capture process

Fig. 2 Sketch diagram of CO$_2$ frost formation process. (1) Crystal growth period. (2) Frost growth period. (3) Frost densification period (Cui et al., 2011)

Fig. 3 The growth process of the CO$_2$ frost layer on the heat exchanger of SC-2. The difference in color represents the direction of the increase in density of the frost layer (from blue to red)

Fig. 4 The structure (a) and dimension (b) of the heat exchanger for capturing CO$_2$
Fig. 5 CO₂ frosting process: (a) no frost; (b) initial frost; (c) growth of the frost; (d) full growth.

Fig. 6 Effect of the flow rate on thickness of the CO₂ frost layer. ($T_{SC,1}$= -30 °C; $T_{SC,2}$= -120 °C; $T_{SC,3}$= -120 °C; $t_1$=240 min; $t_2$=60 min)

Fig. 7 Effect of the flow rate on thermal conductivity ($\lambda_f$) of the CO₂ frost layer. ($T_{SC,1}$= -30 °C; $T_{SC,2}$= -120 °C; $T_{SC,3}$= -120 °C; $t_1$=240 min; $t_2$=60 min)

Fig. 8 Effect of SC-1 on the temperature variation of the SC-2’s cold head. ($T_{SC,1}$= -30 °C; $T_{SC,2}$= -120 °C; $T_{SC,3}$= -120 °C; $v$=2 L/min; $t_1$=240 min; $t_2$=60 min)

Fig. 9 Effect of flow rate on the temperature variation of SC-2’s cold head. ($T_{SC,1}$= -30 °C; $T_{SC,2}$= -120 °C; $T_{SC,3}$= -120 °C; $t_1$=240 min; $t_2$=60 min)

Fig. 10 Effect of temperature variation of SC-1 on the thickness of the frost layer. ($T_{SC,2}$= -120 °C; $T_{SC,3}$= -120 °C; $v$=2 L/min; $t_1$=240 min; $t_2$=60 min)

Fig. 11 Effect of temperature variation of SC-1 on the thermal conductivity of the frost layer.

($T_{SC,2}$= -120 °C; $T_{SC,3}$= -120 °C; $v$=2 L/min; $t_1$=240 min; $t_2$=60 min)
Fig. 1 Schematic diagram of the SC based cryogenic CO\textsubscript{2} capture process

**Flow Direction**
- Red: Flue Gas
- Blue: CO\textsubscript{2}
- Orange: Water
- Dashed: Non-greenhouse Gas

**Abbreviation**
- FGD ---- Flue Gas Desulphurization
- ESP ---- Electrostatic Precipitators
- SCR ---- Selective Catalytic Reduction
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Fig. 7 Effect of the flow rate on thermal conductivity ($\lambda_f$) of the CO$_2$ frost layer. ($T_{SC1}$= -30 °C; $T_{SC2}$= -120 °C; $T_{SC3}$= -120 °C; $t_1$=240 min; $t_2$=60 min)
Fig. 8 Effect of SC-1 on the temperature variation of the SC-2’s cold head. ($T_{SC,1} = -30 \, ^\circ C$; $T_{SC,2} = -120 \, ^\circ C$; $T_{SC,3} = -120 \, ^\circ C$; $v=2$ L/min; $t_1=240$ min; $t_2=60$ min)
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Fig. 10 Effect of temperature variation of SC-1 on the thickness of the frost layer. (T_{SC-2} = -120 °C; T_{SC-3} = -120 °C; \nu = 2 \text{ L/min}; t_1 = 240 \text{ min}; t_2 = 60 \text{ min})
Fig. 11 Effect of temperature variation of SC-1 on the thermal conductivity of the frost layer.

\[ (T_{SC-2} = -120 \degree C; \ T_{SC-3} = -120 \degree C; \ \nu = 2 \ L/min; \ t_1 = 240 \ min; \ t_2 = 60 \ min) \]
Table Captions:

**Table 1** Calculation models for the density of the frost layer.

**Table 2** Empirical models for the thickness of the frost layer.

**Table 3** Correlation models for the thermal conductivity of the frost layer.

**Table 1**

Calculation models for the density of the frost layer.

1) Average frost density calculated by Ogunbameru et al., 1973:

\[ \rho_f = - \ln\left( \frac{G}{I_0} / I \right) / \mu_m \cdot x \]

In which \( G \) is the dimensionless geometrical correction factor. \( I \) is number of \( \gamma \)-radiation counts. \( I_0 \) is number of counts without frost deposit. \( \mu_m \) is mass attenuation factor. \( x \) is frost layer thickness.

2) Frost density is determined from empirical correlation by Hayashi et al., 1977:

\[ \rho_f \bigg|_{\text{emp}} = 650 \cdot \exp \left[ 0.277 \cdot T_f (t) \right] \]

3) Calculation of frost density proposed by Cui et al. in 2011:

\[ \rho_f (\tilde{x}, t) = \rho_g \cdot \left[ 1 - \alpha_i \cdot (\tilde{x}, t) \right] + \rho_i \cdot \alpha_i \cdot (\tilde{x}, t) \]

Where \( \rho_g \) and \( \rho_i \) are the density of gas phase and crystal, respectively. \( \alpha_i \) is the volume fraction of frosted droplets.
4) The local density of the frost layer (Na and Webb, 2004b; Lenic et al., 2009):

$$\rho_f = \varepsilon \rho_s + (1-\varepsilon) \rho_g$$

Here $\varepsilon$ is the local porosity of the frost layer. It is defined as following:

$$\varepsilon = \frac{\rho_s - \rho_f}{\rho_s - \rho_g}, \quad 0 \leq \varepsilon \leq 1$$

In which, $\rho_g$, $\rho_s$ and $\rho_f$ are the density of gas phase, solid phase and frost layer, respectively.

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**Table 2**

Empirical models for the thickness of the frost layer.

1) The frost thickness for each time interval is given by Yun et al., 2002:

$$\delta_j|_{t_{n+1}} = \delta_j + \Delta \delta_j$$

2) The average frost thickness calculated by Cui et al., 2011:

$$\delta_j(t) = \frac{\sum \delta_j(t) \Delta x}{L}$$

In which $\delta_j(t)$ is the local thickness of frost layer. $L$ is the length of the computational domain.

3) The average layer thickness defined by Lenic et al., 2009:

$$\delta_f = \frac{\sum_{k=1}^{M} \delta_{L,k}}{M}$$

Where $\delta_{L,k}$ is the local thickness of frost layer. $M$ is the number of control volumes in horizontal direction.
Table 3

Correlation models for the thermal conductivity of the frost layer.

1) The total thermal conductivity is the sum of surface and inner layer. (Yun et al., 2002):

\[ \lambda_{total} = \frac{\lambda_{f,s} \cdot \delta_{f,s} + \lambda_{f,in} \cdot \delta_{f,in}}{\lambda_{f,in} \cdot \delta_{f,s} + \lambda_{f,s} \cdot \delta_{f,in}} \]

Here, the thermal conductivity of surface is given by:

\[ \lambda_{f,s} = \lambda_a + \rho_{a,s} \cdot c_{p,a,s} \cdot u_f^2 \cdot C_f / 2 \cdot \left[ y(t) + \delta_b \right] / 2.5 \]

the thermal conductivity of inner layer is calculated by Woodside in 1958:

\[ \lambda_{f,in} = \lambda_a / (1 - 6 \cdot s / \pi)^{1/3} \cdot \left[ 1 - \left( c^2 - 1 \right) / 2 \cdot \ln \left( (c+1) / (c-1) \right) \right] \]

2) The frost thermal conductivity is related to temperature as a sole parameter (Sahin, 2000):

\[ \lambda_f(T) = 0.131 \times 10^{-6} (1 - \beta) \cdot \frac{h_s \cdot P_{g \infty} \cdot P_g}{T_s^{1.84} \cdot R^3 \cdot T_s^{1.06}} \cdot \exp \left[ \frac{h_s}{R} \cdot \left( \frac{1}{T_s} - \frac{1}{T_i} \right) \right] + 1.202 \times 10^{-3} \cdot \beta \cdot (c \cdot T_s - 1) \cdot (c - 1)^{1.0465 + 0.017T_s} \times 10^{-5} \]

Where \( h_s \) is the latent heat of sublimation. \( P_g \) is the partial pressure of specific gas phase at initial temperature \( T_s \). \( T_i \) is the temperature of frost surface. \( \beta \) is the volumetric ratio of frost deposition. \( \rho_c \) is the density of crystal.

3) The empirical equation validated by Yonko and Sepsy, 1967:

\[ \lambda_f = 0.024248 + 0.00072311 \rho_f + 0.000001183 \rho_f^2 \]

In which \( \rho_f \) is the density of frost layer.

4) The expression demonstrated by Sanders, 1974:
\[ \lambda_f = 1.202 \times 10^{-3} \cdot \rho_f^{0.961} \]

5) The correlation proposed by Lee et al., 2003:

\[ \lambda_f = 0.132 + 3.13 \times 10^{-1} \cdot \rho_f + 1.6 \times 10^{-7} \cdot \rho_f^2 \]