

Luminescent Ir(III) complexes containing
benzothiazole-based tridentate ligands:
synthesis, characterization, and application
to organic light-emitting diodes

著者別名	桑原 純平, 神原 貴樹
journal or publication title	Dalton transactions
volume	41
number	1
page range	44-46
year	2012-01
権利	(C) Royal Society of Chemistry 2012
URL	http://hdl.handle.net/2241/117672

doi: 10.1039/C1DT11560J

Cite this: DOI: 10.1039/c0xx00000x

www.rsc.org/xxxxxx

COMMUNICATION

Luminescent Ir(III) complexes containing benzothiazole-based tridentate ligands: synthesis, characterization, and application to organic light-emitting diodes

Junpei Kuwabara,^a Tomomi Namekawa,^a Masa-aki Haga,^b Takaki Kanbara^{*a}

⁵ Received (in XXX, XXX) Xth XXXXXXXXXX 20XX, Accepted Xth XXXXXXXXXX 20XX

DOI: 10.1039/b000000x

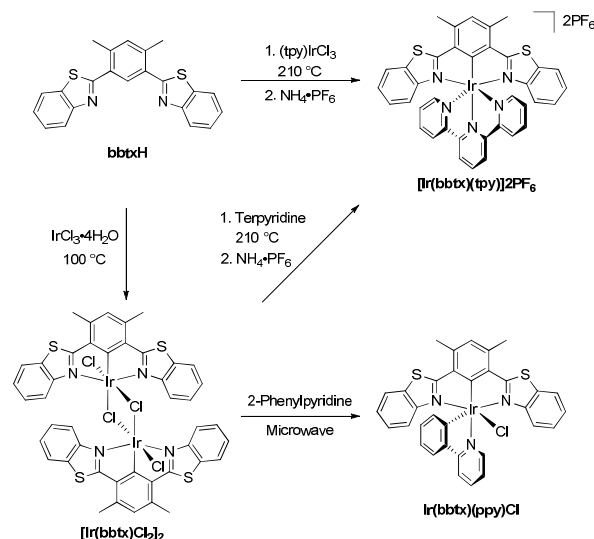
Ir(III) complexes that contain benzothiazole-based tridentate ligands were synthesized and their crystal structures and luminescent properties were examined. A neutral complex had a high quantum yield (89%) and performed well as an emissive material for organic light-emitting diodes.

Keywords: Ir complex / Pincer ligand / phosphorescence / OLEDs

Iridium complexes composed of cyclometallated ligands are promising phosphorescent materials for organic light-emitting diodes (OLEDs).¹ Anionic bidentate ligands such as 2-phenylpyridine (ppy) and 2-phenylbenzothiazole (bt) are often found in the structures of emissive Ir complexes as cyclometallated ligands.² In particular, Ir(ppy)₃ is used widely in OLEDs owing to its high quantum yield and thermal stability. In general, tridentate cyclometallated ligands, also called pincer ligands, impart better thermal stability to complexes when bound with the metal center than do bidentate ligands.³ Therefore, Ir complexes composed of pincer ligands are believed to exhibit excellent thermal stability, which is beneficial for fabricating OLEDs and the stability of the devices that they are included in. Williams and Haga reported independently on Ir complexes composed of pincer ligands and their strong emission properties.⁴ Although their reports alluded to the great potential of pincer Ir complexes as emissive materials, there are only limited examples of pincer Ir complexes being applied to manufacturing OLEDs.⁵ Ir complexes that contain 2-phenylbenzothiazole have good emissive performances, and therefore, tridentate ligands that are composed of two benzothiazole units are promising pincer ligand candidates. The pincer ligand was synthesized originally as a ligand of a Pt complex in our previous study.⁶ We herein report on the synthesis of Ir complexes composed of benzothiazole-based pincer ligands and their performance as emitting materials

in OLEDs.

1,3-Bis(2-benzothiazolyl)-4,6-dimethylbenzene (**bbtxH**) was synthesized from 1,3-diiodo-4,6-dimethylbenzene by the Pd-catalyzed direct arylation of benzothiazole in 69% yield.⁷ To check the reactivity of the ligand, the reaction between (tpy)IrCl₃ (tpy = terpyridine) and **bbtxH** at 210 °C in ethylene glycol was examined (Scheme 1). After treatment using NH₄PF₆, dicationic [Ir(**bbtx**)(tpy)]2PF₆ was obtained in 66% yield through purification by crystallization. Although the reaction resulted in a single product, the combination of (tpy)IrCl₃ and 1,3-bis(2-benzothiazolyl)-5-methylbenzene⁶ gave a mixture of products, presumably due to undesired cyclometallating reactions at positions such as 4 or 6 on the central benzene moiety. These results indicate that two methyl groups at 4 and 6 positions on **bbtxH** play important roles in the selective cyclometallation at the 2 position. The reaction of **bbtxH** and IrCl₃·4H₂O gave the Cl-bridging dimer complex [Ir(**bbtx**)Cl₂]₂ in 89% yield (Scheme 1). Since [Ir(**bbtx**)Cl₂]₂ is insoluble in commonly used organic solvents such as CHCl₃, THF, DMF, and DMSO, the complex was characterized by mass spectrometry and elemental analysis. [Ir(**bbtx**)(tpy)]2PF₆ was also obtained from [Ir(**bbtx**)Cl₂]₂ and terpyridine in 64% yield. The reaction of [Ir(**bbtx**)Cl₂]₂ with 2-



Scheme 1. Syntheses of Ir complexes with a benzothiazole-based tridentate ligand.

^aTsukuba Research Center for Interdisciplinary Materials Science (TIMS), Graduate School of Pure and Applied Sciences, University of Tsukuba, 1-1-1, Tennodai, Tsukuba, Japan; E-mail: kanbara@ims.tsukuba.ac.jp

^bDepartment of Applied Chemistry, Faculty of Science and Engineering, Chuo University, 1-13-27, Kasuga, Bunkyo-ku, Tokyo, Japan.

† Electronic Supplementary Information (ESI) available: [details of any supplementary information available should be included here]. See DOI: 10.1039/b000000x/

phenylpyridine was carried out by applying microwave radiation for 4 min; it resulted in a neutral complex, **Ir(bbtx)(ppy)Cl**, in 76% yield. Applying microwave radiation was better than employing conventional heating in terms of the yield, at 76% versus 32%, and a reaction time of 4 min versus 20 h, respectively.⁸

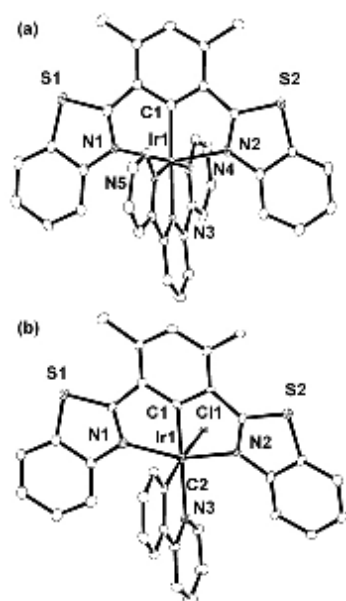


Fig. 1. ORTEP drawings of (a) **[Ir(bbtx)(tpy)]₂PF₆** and (b) **Ir(bbtx)(ppy)Cl** with thermal ellipsoids shown at the 30% probability level. Hydrogen atoms, counter anions, and solvating molecules are omitted for clarity.

The solid-state structures of **[Ir(bbtx)(tpy)]₂PF₆** and **Ir(bbtx)(ppy)Cl** were determined by X-ray diffraction studies (Fig. 1).^{9,10} Both complexes possessed octahedral geometry around the Ir center. In **Ir(bbtx)(ppy)Cl**, the cyclometalated carbon of ppy was located at the trans position of the Cl ligand, which is consistent with the reported structure of the Ir pincer complex.⁴

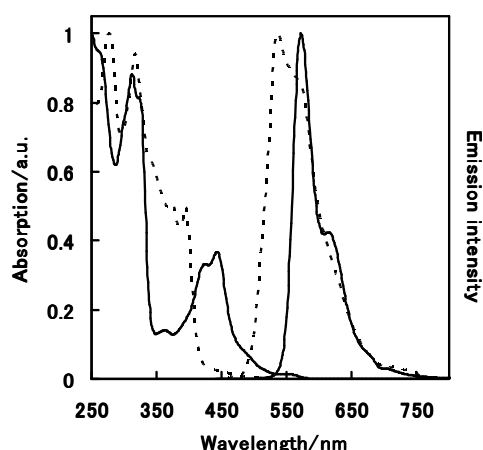


Fig. 2. UV/Vis absorption and photoluminescence spectra of **[Ir(bbtx)(tpy)]₂PF₆** in CH₃CN at r.t. (dashed line) and **Ir(bbtx)(ppy)Cl** in CH₂Cl₂ at r.t. (solid line).

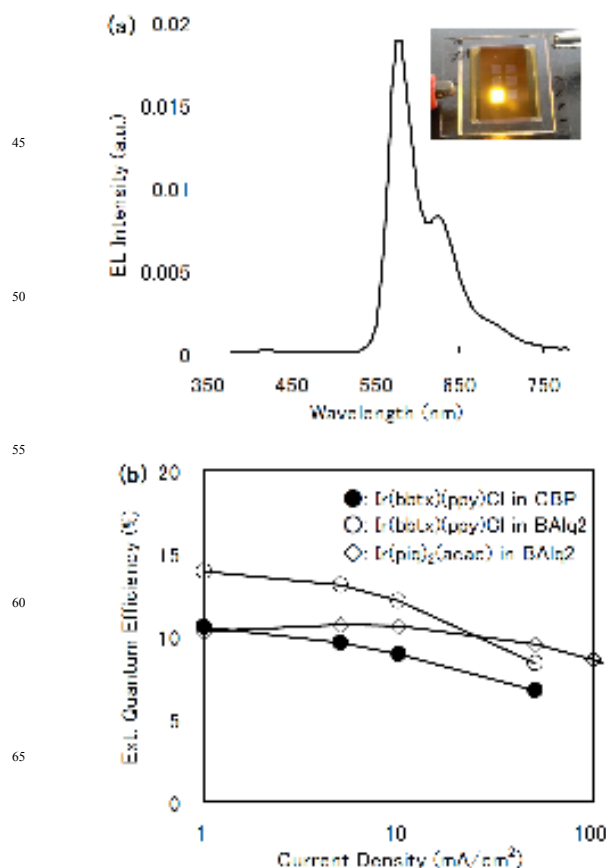


Fig. 3. (a) Electroluminescence (EL) spectrum of device I (inset: photograph of EL of the device at 3 V). (b) External quantum efficiencies of devices I, II, and III.

The UV/Vis absorption and photoluminescence spectra of **[Ir(bbtx)(tpy)]₂PF₆** and **Ir(bbtx)(ppy)Cl** at r.t. are shown in Fig. 2. **Ir(bbtx)(ppy)Cl** exhibited characteristic absorption at around 450 nm, and this absorption was absent for **[Ir(bbtx)(tpy)]₂PF₆**. The absorption was assigned tentatively to metal-to-ligand charge transfer (MLCT) on the basis of comparisons with Ir complexes composed of NCN ligands.⁴ **Ir(bbtx)(ppy)Cl** exhibited a strong emission at 572 nm when excited at 443 nm. The emission spectrum had a vibrational structure with a small side band at 1223 cm⁻¹. The quantum yield of **Ir(bbtx)(ppy)Cl** was as high as 89% in degassed CH₂Cl₂ at r.t. The emission lifetime of **Ir(bbtx)(ppy)Cl** was 1.1 × 10⁻⁶ s. Since the quantum yield was comparable to Ir(ppy)₃,¹¹ **Ir(bbtx)(ppy)Cl** was expected to have high potential for use as an emissive material in OLEDs. On the other hand, **[Ir(bbtx)(tpy)]₂PF₆** exhibited a weak emission at 537 nm, where the quantum yield was less than 0.1% even in degassed CH₃CN. The low quantum yield of **[Ir(bbtx)(tpy)]₂PF₆** is presumably due to low energy level of a d orbital on the Ir center, which inhibits MLCT leading to effective emission.⁴ⁱ

The thermal stability of **Ir(bbtx)(ppy)Cl** was evaluated by thermogravimetric analysis. The 5% decomposition temperature of **Ir(bbtx)(ppy)Cl** was 480 °C, which was higher than that of Ir(ppy)₃ at 413 °C.^{2c} This result proved to be the advantage of the pincer ligand in terms of thermal stability. Owing to its high

thermal stability and high fluorescence quantum yield, **Ir(bbtX)(ppy)Cl** was evaluated as a candidate dopant for the emitting layer of OLEDs. To examine the electrophosphorescent properties of **Ir(bbtX)(ppy)Cl**, three OLEDs (I, II, and III) were fabricated with different host materials and dopants. In the emission layer of device I, 4,4'-bis(9-dicarbazolyl)-2,2'-biphenyl (CBP) was used as the host material and **Ir(bbtX)(ppy)Cl** was used as the dopant. In device II, the host material was bis(2-methyl-8-quinolinolato)phenolatealuminium(III) (BALq2) instead of CBP and the dopant was **Ir(bbtX)(ppy)Cl**. In device III, the host material was BALq2 and the dopant was bis(1-phenylisoquinoline)(acetylacetonate)iridium(III), Ir(piq)₂(acac). Fig. 3a shows the electroluminescence spectrum of device I, which is similar to its photoluminescence spectrum in the diluted solution state. This result indicates that **Ir(bbtX)(ppy)Cl** was well dispersed throughout CBP. The coordinates of the CIE chromaticity of device I were $x = 0.55$ and $y = 0.44$ at 1 mA/cm². The external quantum efficiency of device I was 10.5% at 1 mA/cm² (Fig. 3b). Since device II possessed a higher external quantum efficiency (14.0%) than device I, BALq2 was a suitable host material for **Ir(bbtX)(ppy)Cl**. In addition, device II had the advantage of a longer lifetime of electroluminescence than device I (Fig. S-2).† Device II had an external quantum efficiency comparable to device III with Ir(piq)₂(acac), which is representative of a red emission dopant.¹²

In summary, the neutral Ir complex that contains the benzothiazole-based pincer ligand has a higher quantum yield and thermal stability than Ir complexes bearing bidentate ligands. The Ir complex can serve as an efficient emissive dopant in OLEDs. This investigation was an important step for the development of pincer Ir complexes as emissive materials for OLEDs. Further investigations such as optimizing devices that contain **Ir(bbtX)(ppy)Cl** as a dopant and developing new designs of pincer ligands are currently underway.

Acknowledgements

We are grateful to the Chemical Analysis Center of the University of Tsukuba for X-ray diffraction studies, elemental analyses, and NMR spectroscopy in this research. We thank Prof. T. Nabeshima, Prof. M. Yamamura, Prof. Y. Nishimura, Dr. T. Inoue, and S. Kitagawa for assistance in measuring the quantum yields, emission lifetime, and electrophosphorescent properties of the complexes.

Notes and references

- (a) M. A. Baldo, S. Lamansky, P. E. Burrows, M. E. Thompson and S. R. Forrest, *Appl. Phys. Lett.*, 1999, **75**, 4; (b) M. Ikai, S. Tokito, Y. Sakamoto, T. Suzuki and Y. Taga, *Appl. Phys. Lett.*, 2001, **79**, 156; (c) C. Adachi, M. A. Baldo, M. E. Thompson and S. R. Forrest, *J. Appl. Phys.*, 2001, **90**, 5048; (d) M. S. Lowry and S. Bernhard, *Chem. Eur. J.*, 2006, **12**, 7970; (e) C. Ulbricht, B. Beyer, C. Friebe, A. Winter and U. S. Schubert, *Adv. Mater.*, 2009, **21**, 4418; (f) Y. You and S. Y. Park, *Dalton Trans.*, 2009, 1267; (g) Y. Chi and P.-T. Chou, *Chem. Soc. Rev.*, 2010, **39**, 638.
- (a) S. Lamansky, P. Djurovich, D. Murphy, F. Abdel-Razzaq, R. Kwong, I. Tsyba, M. Bortz, B. Mui, R. Bau and M. E. Thompson, *Inorg. Chem.*, 2001, **40**, 1704; (b) S. Lamansky, P. Djurovich, D. Murphy, F. Abdel-Razzaq, H.-E. Lee, C. Adachi, P. E. Burrows, S. R. Forrest and M. E. Thompson, *J. Am. Chem. Soc.*, 2001, **123**, 4304; (c) A. Tsuboyama, H. Iwawaki, M. Furugori, T. Mukaide, J. Kamatani, S. Igawa, T.

- Moriyama, S. Miura, T. Takiguchi, S. Okada, M. Hoshino and K. Ueno, *J. Am. Chem. Soc.*, 2003, **125**, 12971; (d) J.-Y. Hung, Y. Chi, I.-H. Pai, Y.-C. Yu, G.-H. Lee, P.-T. Chou, K.-T. Wong, C.-C. Chen and C.-C. Wu, *Dalton Trans.*, 2009, 6472; (e) K. Tsuchiya, S. Yagai, A. Kitamura, T. Karatsu, K. Endo, J. Mizukami, S. Akiyama and M. Yabe, *Eur. J. Inorg. Chem.*, 2010, 926; (f) V. K. Rai, M. Nishiura, M. Takimoto and Z. Hou, *Chem. Commun.*, 2011, **47**, 5726.
- M. Albrecht and G. van Koten, *Angew. Chem., Int. Ed.*, 2001, **40**, 3750.
- (a) A. J. Wilkinson, A. E. Goeta, C. E. Foster and J. A. G. Williams, *Inorg. Chem.*, 2004, **43**, 6513; (b) A. J. Wilkinson, H. Puschmann, J. A. K. Howard, C. E. Foster and J. A. G. Williams, *Inorg. Chem.*, 2006, **45**, 8685; (c) V. L. Whittle and J. A. G. Williams, *Inorg. Chem.*, 2008, **47**, 6596; (d) T. Yutaka, S. Obara, S. Ogawa, K. Nozaki, N. Ikeda, T. Ohno, Y. Ishii, K. Sakai and M. Haga, *Inorg. Chem.*, 2005, **44**, 4737; (e) S. Obara, M. Itabashi, F. Okuda, S. Tamaki, Y. Tanabe, Y. Ishii, K. Nozaki and M. Haga, *Inorg. Chem.*, 2006, **45**, 8907; (f) L. Yang, F. Okuda, K. Kobayashi, K. Nozaki, Y. Tanabe, Y. Ishii and M. Haga, *Inorg. Chem.*, 2008, **47**, 7154; (g) M. Ashizawa, L. Yang, K. Kobayashi, H. Sato, A. Yamagishi, F. Okuda, T. Harada, R. Kuroda and M. Haga, *Dalton Trans.*, 2009, 1700; (h) J. A. G. Williams, A. J. Wilkinson and V. L. Whittle, *Dalton Trans.*, 2008, 2081; (i) J. A. G. Williams, *Chem. Soc. Rev.*, 2009, **38**, 1783.
- D. Choi, T. Kim, S. M. Reddy and J. Kang, *Inorg. Chem. Commun.*, 2009, **12**, 41.
- K. Okamoto, T. Kanbara, T. Yamamoto and A. Wada, *Organometallics*, 2006, **25**, 4026.
- (a) A. Yokooji, T. Okazawa, T. Satoh, M. Miura and M. Nomura, *Tetrahedron*, 2003, **59**, 5685; (b) D. Alagille, R. M. Baldwin and G. D. Tamagnan, *Tetrahedron Lett.*, 2005, **46**, 1349; (c) A. Mori, A. Sekiguchi, K. Masui, T. Shimada, M. Horie, K. Osakada, M. Kawamoto, T. Ikeda, *J. Am. Chem. Soc.*, 2003, **125**, 1700.
- H. Konno and Y. Sasaki, *Chem. Lett.*, 2003, **32**, 252.
- Crystallographic details for **[Ir(bbtX)(tpy)]2PF₆·2DMF**; C₄₃H₄₀N₇F₁₂IrO₂P₂S₂, $M = 1233.1$, Rigaku RAXIS-RAPID (Mo $K\alpha$ radiation), $T = 88$ K, Triclinic spacegroup $P-1$, $a = 8.8979(3)$ Å, $b = 14.1372(5)$ Å, $c = 19.0729(6)$ Å, $\alpha = 97.1587(12)$, $\beta = 90.6934(11)$, $\gamma = 110.3673(12)^\circ$, $V = 2227.70(13)$ Å³, $Z = 2$, $D_{\text{calc}} = 1.838$ g/cm³, 21949 measured reflections, 10019 independent reflections [$R_{\text{int}} = 0.079$], $R_1[\text{I} > 2\sigma(\text{I})] = 0.0404$, $wR_2(\text{all data}) = 0.0969$, GOF = 1.078. CCDC 834013. Selected bond distances of **[Ir(bbtX)(tpy)]2PF₆** (Å): Ir-C1 = 1.972(4), Ir-N1 = 2.093(3), Ir-N2 = 2.087(4), Ir-N3 = 2.030(3), Ir-N4 = 2.066(3), Ir-N5 = 2.056(3).
- Crystallographic details for **[Ir(bbtX)(ppy)Cl]·2CHCl₃**; C₃₅H₂₅N₃IrS₂Cl₇, $M = 992.1$, Rigaku RAXIS-RAPID (Mo $K\alpha$ radiation), $T = 88$ K, Triclinic spacegroup $P-1$, $a = 9.7217(5)$ Å, $b = 12.3438(6)$ Å, $c = 16.1530(8)$ Å, $\alpha = 79.6257(15)$, $\beta = 87.9281(17)$, $\gamma = 71.2491(15)^\circ$, $V = 1805.03(16)$ Å³, $Z = 2$, $D_{\text{calc}} = 1.825$ g/cm³, 17542 measured reflections, 8161 independent reflections [$R_{\text{int}} = 0.098$], $R_1[\text{I} > 2\sigma(\text{I})] = 0.0563$, $wR_2(\text{all data}) = 0.1348$, GOF = 0.943. CCDC 834012. Selected bond distances (Å): Ir-C1 = 1.956(6), Ir-C2 = 2.013(7), Ir-N1 = 2.073(6), Ir-N2 = 2.083(7), Ir-N3 = 2.158(5), Ir-Cl1 = 2.5034(15).
- (a) S. Sprouse, K. A. King, P. J. Spellane and R. J. Watts, *J. Am. Chem. Soc.*, 1984, **106**, 6647; (b) A. Endo, K. Suzuki, T. Yoshihara, S. Tobita, M. Yahiro and C. Adachi, *Chem. Phys. Lett.*, 2008, **460**, 155.
- Y.-J. Su, H.-L. Huang, C.-L. Li, C.-H. Chien, Y.-T. Tao, P.-T. Chou, S. Datta and R.-S. Liu, *Adv. Mater.*, 2003, **15**, 884.